

 Estd. 1962 "A++" Accredited by NAAC (2021) With CGPA 3.52	SHIVAJI UNIVERSITY, KOLHAPUR - 416004, MAHARASHTRA PHONE:EPABX-2609000, www.unishivaji.ac.in , bos@unishivaji.ac.in शिवाजी विद्यापीठ, कोल्हापूर - ४१६००४, महाराष्ट्र दूरध्वनी-ईपीएबीएक्स -२६०९०००, अभ्यासमंडळे विभाग दूरध्वनी ०२३१-२६०९०९४ ०२३१-२६०९४८७		
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SU/BOS/Science/06

Date: 01/01/2024

To,

The Principal, All Concerned Affiliated Colleges/Institutions Shivaji University, Kolhapur	The Head/Co-ordinator/Director All Concerned Department (Science) Shivaji University, Kolhapur.
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Subject: Regarding syllabi of B.Sc. Part-III (Sem. V & VI) as per NEP-2020 (1.0) degree programme under the Faculty of Science and Technology.

Sir/Madam,

With reference to the subject mentioned above, I am directed to inform you that the university authorities have accepted and granted approval to the revised syllabi, nature of question paper and equivalence of B.Sc. Part-III (Sem. V & VI) as per NEP-2020 (1.0) degree programme under the Faculty of Science and Technology.

B.Sc.-III (Sem. V & VI) as per NEP-2020 (1.0)			
1.	Mathematics	12.	Computer Science (Opt)
2.	Statistics	13.	Computer Science (Entire)
3.	Physics	14.	Information Technology (Entire)
4.	Microbiology	15.	Food Science and Technology (Entire)
5.	Industrial Microbiology	16.	Food Science
6.	Electronics	17.	Food Science and Quality Control (Entire)
7.	Chemistry	18.	Food Technology & Management (Entire)
8.	Sugar Technology (Entire)	19.	Biochemistry
9.	Geology	20.	Biotechnology (Optional/Vocational)
10.	Zoology	21.	Biotechnology (Entire)
11.	Botany	22.	Environmental Science (Entire)

This syllabus, nature of question and equivalence shall be implemented from the academic year 2024-2025 onwards. A soft copy containing the syllabus is attached herewith and it is also available on university website www.unishivaji.ac.in NEP-2020(Online Syllabus)

The question papers on the pre-revised syllabi of above-mentioned course will be set for the examinations to be held in October /November 2024 & March/April 2025. These chances are available for repeater students, if any.

You are, therefore, requested to bring this to the notice of all students and teachers concerned.

Thanking you,


 By Registrar
 Dr. S. M. Kubal

Copy to:

Shivaji University, Kolhapur



Accredited By NAAC with 'A++' Grade

Syllabus for

Bachelor of Science Part-III

Chemistry

(NEP-1.0)

To be implemented from

June, 2024 onwards

INTRODUCTION

This syllabus is prepared to give the sound knowledge and understanding of chemistry to undergraduate students at last year of the B.Sc. degree course. The goal of the syllabus is to make the study of chemistry as stimulating, interesting and relevant as possible. The syllabus is prepared by keeping in mind the aim to make students capable of studying chemistry in academic and industrial courses and to expose the students, to develop interest in them in various fields of chemistry. The new and updated syllabus is based on disciplinary approach with vigour and depth taking care the syllabus is not heavy at the same time it is comparable to the syllabi of other universities at the same level. The syllabus is prepared after discussions of number of faculty members of the subject and by considering the existing syllabi of B.Sc. Part-I, II & III, new syllabi of XIth & XIIth standards, syllabi of NET and SET exams. U.G.C. model curriculum, syllabi of different entrance examination and syllabi of other Universities. The units of the syllabus are well defined and the scope is given in detail. The periods required for units are given. The lists of reference books are given in detail.

OBJECTIVES

1. To promote understanding of basic facts and concepts in Chemistry while retaining the excitement of Chemistry
2. To make students capable of studying Chemistry in academic and Industrial courses and to expose the students to different processes used in Industries and their applications.
3. To expose the students to various emerging new areas of Chemistry and apprise them with their prevalent in their future studies and their applications in various spheres of chemical sciences.
4. To develop problem solving skills in students.
5. To developed ability and to acquire the knowledge of terms, facts, concepts, processes, techniques and principles of subjects.
6. To develop ability to apply the knowledge of contents of principles of chemistry.
7. To inquire of new knowledge of chemistry and developments therein.
8. To expose and to develop interest in the fields of chemistry
9. To develop proper aptitude towards the subjects
10. To develop the power of appreciations, the achievements in Chemistry and role in nature and society.
11. To develop skills required in chemistry such as the proper handling of apparatus and Chemicals.

Shivaji University, Kolhapur
B.O.S. in Chemistry
B. Sc. Part III (NEP 1. 0) SEMESTER-V Syllabus

To be implemented from June – 2024

Equivalence

Sr. No.	Title of old paper (Syllabus implemented from June-2020)	Title of new paper (To be implemented from June-2024)
1	Paper IX DSE-E5 and XIII DSE-F5 Inorganic Chemistry	Paper IX DSE-E5 and XIII DSE-F5 Inorganic Chemistry
2	Paper- X DSE-E6 and XIV DSE-F6 Organic Chemistry	Paper- X DSE-E6 and XIV DSE-F6 Organic Chemistry
3	Paper XI DSE-E7 and XV- DSE-F7 Physical Chemistry	Paper XI DSE-E7 and XV- DSE-F7 Physical Chemistry
4	Paper XII-DSE-E8 and XVI DSE-F8: Analytical Chemistry	Paper XII-DSE-E8 and XVI DSE-F8: Analytical Chemistry

A repeater candidate, if any, will be allowed to appear for Theory and Practical examination as per old course up to March / April 2025 examination.

B. Sc. Part - III (NEP- 1.0) SEMESTER-V

Paper No. DSE-E5, Chemistry Paper No.– IX (Inorganic Chemistry)

(Theory Credits: 02, 30 hours.)

Expected Learning Outcomes:

Name of the topic	Expected Learning Outcome
1. Acids bases and Non-aqueous solvents	Useful for the study of role of acids and bases in Chemistry. The study of non –aqueous solvents is important to learn all chemical properties of solutes and from the research Point of view.
2. Metal ligand bonding in transition metal complexes	Useful to understand geometry, stability and Nature of bonding between metal ion and ligand in complexes.
3. Metals, semiconductors and Superconductors	The topic deals with the synthesis and the applications of the semiconductors and Superconductors in electrical and electronic devices.
4. Organometallic compounds	The structure, method of preparation and the Applications of organometallic compound in various fields are explained.
5. Catalysis	The classification, types, mechanism and applications of catalyst in industrial fields is explained.

Unit-1. Acids, Bases and Non-aqueous Solvents

[6]

1.1 Introduction to theories of Acids and Bases – Arrhenius concept, Bronsted-Lowry concept, Lewis Concept, Lux-Flood Concept (definition and examples)

1.2 Hard and Soft Acids and Bases.(HSAB Concept)

1.2.1 Classification of acids and bases as hard, soft and borderline.

1.2.2 Pearson’s HSAB concept.

1.2.3 Acid–Base strength and hardness-softness.

1.2.4 Applications and limitations of HSAB principle.

1.3 Chemistry of Non-aqueous Solvents.

1.3.1 Introduction, definition and characteristics of solvents.

1.3.2 Classification of solvents.

- 1.3.3 Physical properties and Acid-Base reactions in Liquid Ammonia (NH_3) and Liquid Sulphur Dioxide (SO_2).

Unit-2. Metal Ligand bonding in Transition Metal Complexes [8]

2.1 Crystal field theory (CFT)

2.2 Introduction: Shapes of d-orbital's, Basic assumptions of CFT.

2.2.1 Crystal field splitting of d-orbital's of metal ion in octahedral, tetrahedral, square planar complexes and John-Teller distortion.

2.2.2 Factors affecting the Crystal field splitting.

2.2.3 High spin and low spin octahedral complexes w. r. t. Co(III) .

2.2.4 Crystal Field stabilization energy (CFSE), Calculation with respect to octahedral complexes only.

2.2.5 Limitations of CFT.

2.3 Molecular orbital theory (MOT).

2.3.1 Introduction.

2.3.2 MOT of octahedral complexes with sigma bonding such as $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$, $[\text{CoF}_6]^{3-}$, $[\text{Co}(\text{NH}_3)_6]^{3+}$.

2.3.3 Merits and demerits of MOT.

Unit-3. Metals, Semiconductors and Superconductors. [8]

3.1 Introduction.

3.1 Properties of metallic solids.

3.2 Theories of bonding in metal.

3.2.1 Free electron theory.

3.2.2 Molecular orbital theory(Band theory).

3.3 Classification of solids as conductor, insulators and semiconductors on the basis of band theory.

3.4 Semiconductors-Types-intrinsic and extrinsic and applications of semiconductors.

3.5 Superconductors: Ceramic superconductors- Preparation and structures of mixed oxide $\text{YBa}_2\text{Cu}_3\text{O}_{7-x}$.

3.6 Applications of superconductors.

Unit-4. Organometallic Chemistry.

[4]

- 4.1 Definition, Nomenclature of organometallic compounds.
- 4.2 Synthesis and structural study of alkyl and aryl compounds of Be and Al.
- 4.3 Mononuclear carbonyls – Nature of bonding in simple mononuclear carbonyls.: $[\text{Ni}(\text{CO})_4]$, $[\text{Fe}(\text{CO})_5]$, $[\text{Cr}(\text{CO})_6]$.

Unit-5. Catalysis

[4]

- 5.1 Introduction
- 5.2 Classification of catalytic reaction- Homogenous and Heterogeneous
- 5.3 Types of Catalysis.
- 5.4 Characteristics of catalytic reactions.
- 5.5 Mechanism of catalysis.
 - 5.5.1 Intermediate compound formation theory.
 - 5.5.2 Adsorption theory.
- 5.6 Industrial applications of catalysis.

Reference Books:

1. Concise Inorganic Chemistry (ELBS, 5th Edition) – J.D. Lee.
2. Inorganic Chemistry (ELBS, 3rd Edition) D. F. Shriver, P.W. Atkins, C.H. Langford, Oxford University Press, 2nd Edition.
3. Basic Inorganic Chemistry: Cotton and Wilkinson.
4. Advanced Inorganic Chemistry (4th Edn.) Cotton and Wilkinson.
5. Concepts and Models of Inorganic Chemistry: Douglas and McDaniel. 3rd Edition. John Wiley publication.
6. Structural principles in inorganic compounds. W. E. Addison.
7. Theoretical principles of Inorganic Chemistry – G. S. Manku.
8. Theoretical Inorganic Chemistry by Day and Selbina.
9. Co-ordination compounds. SFA Kettle.
10. Essentials of Nuclear Chemistry by H. J. Arnikar.
11. Nuclear Chemistry by M. N. Sastri.
12. Organometallic Chemistry by R. C. Mahrotra, A. Sing, Wiley Eastern Ltd. New Delhi.
13. Inorganic Chemistry by A. G. Sharpe, Addison-Wisley Longman – Inc.

14. Principles of Inorganic Chemistry by Puri, Sharma and Kalia, Vallabh Publication.
Pitampur Delhi.
15. Text book of Inorganic Chemistry by K. N. Upadhyaya Vikas Publishing House—New
Delhi.
16. Inorganic Chemistry 3rd Edn G. L. Miessler and D. A. Tarr, Pearson publication.
17. Co-ordination compounds by Baselo and Pearson.
18. UGC Inorganic chemistry by H. C. Khera, Pragati prakashan
19. UGC Advanced Inorganic Chemistry by Agarwal and Keemtilal, Pragati Prakashan

B. Sc. Part III (NEP 1. 0) SEMESTER-V

Paper No. DSE-E6 Chemistry Paper No. X (Organic Chemistry)

(Theory Credits: 02, 30 hours.)

Expected Learning Outcomes:

Name of the Topic	Expected Learning Outcomes
1. Introduction to Spectroscopy	Understanding of energy associated with electromagnetic radiation and its use in analytical technique.
2. UV-Vis Spectroscopy	Knowledge of chromophore, auxochrome and calculation of λ_{max} .
3. IR Spectroscopy	Knowledge of vibrational transitions, regions of IR spectrum, functional group recognition.
4. NMR Spectroscopy	Understanding of magnetic, non-magnetic nuclei, shielding-desielding, chemical shift, splitting Pattern
5. Introduction to Mass spectroscopy. Combined Spectroscopic Problems based on UV-Vis, IR and NMR data.	Knowledge of molecular ion, fragmentation pattern and different types of ions produced. Student will predict the structure of organic compound with the help of provided spectral data.

Unit 1. Introduction to Spectroscopy

[3]

- 1.1 Meaning of spectroscopy.
- 1.2 Nature of electromagnetic radiation: wavelength, frequency, energy, amplitude, wave number and their relationship.
- 1.3 Different units of measurement of wavelength and frequency.
- 1.4 Different regions of electromagnetic radiations.
- 1.5 Interaction of radiation with matter: absorption, emission, fluorescence and scattering.
- 1.6 Types of spectroscopy and advantages of spectroscopic methods.
- 1.7 Energy types and energy levels of atoms and molecules.

Unit 2. UV-visible Spectroscopy

[4]

- 2.1 Introduction.
- 2.2 Beer-Lambert's law, absorption of UV radiation by organic molecules leading to different excitations.
- 2.3 Terms used in UV Spectroscopy: Chromophore, Auxochrome, Bathochromic shift, hypsochromic shift, hyperchromic and hypochromic effect.
- 2.4 Modes of electromagnetic transitions.
- 2.5 Effect of conjugation on position of UV band.
- 2.6 Calculation of λ_{max} by Woodward and Fischer rules for dienes and enones.
- 2.7 Colour and visible spectrum.
- 2.8 Applications of UV Spectroscopy.

Unit 3. IR Spectroscopy

[5]

- 3.1 Introduction.
- 3.2 Principle of IR Spectroscopy.
- 3.3 Instrumentation - Schematic diagram.
- 3.4 Fundamental modes of vibrations, types and calculation.
- 3.5 Conditions for absorption of IR radiations.
- 3.6 Regions of IR spectrum, fundamental group region, finger print region.
- 3.7 Hook's Law for Calculation of vibrational frequency.
- 3.8 Factors affecting IR absorption frequency.
- 3.9 Characteristic of IR absorption of following functional groups a) alkanes, alkenes,

alkynes b) alcohol and phenols c) ethers d) carbonyl compounds e) amines f) nitro compounds and g) aromatic compounds.

Unit 4. NMR Spectroscopy

[7]

- 4.1 Introduction.
- 4.2 Principle of PMR Spectroscopy.
- 4.3 NMR- Instrumentation - Schematic diagram.
- 4.4 Magnetic and nonmagnetic nuclei.
- 4.5 Chemical shift: definition, measurement, calculation, Factors affecting Chemical shift.
- 4.6 Shielding & deshielding.
- 4.7 Peak Integration.
- 4.8 Merits of TMS as PMR reference compound.
- 4.9 Coupling Constant.
- 4.10 Types of Coupling Constant.
- 4.11 Spin-spin splitting (n+1 rule).
- 4.12 Applications of NMR Spectroscopy.

Unit 5. Introduction to Mass Spectroscopy.

[6]

- 5.1 Introduction.
- 5.2 Principles of mass spectroscopy.
- 5.3 Mass spectrometer - Schematic diagram.
- 5.4 Types of ions produced during fragmentation.
- 5.5 Nitrogen rule
- 5.6 Fragmentation patterns of: alkanes, alkenes and carbonyl compounds.
- 5.7 McLafferty rearrangement.
- 5.8 Applications of Mass Spectroscopy.

Unit 6. Combined Spectroscopic Problems based on UV, IR and NMR Spectral data. [5]

Reference Books: (Use recent editions)

- 1. Absorption Spectroscopy of Organic Molecules by V. M. Parikh.
- 2. Spectroscopy of Organic compounds by P. S. Kalsi.
- 3. Elementary Organic Absorption Spectroscopy by Y. R. Sharma.

4. Instrumental Methods of Analysis (7th edition) by Willard, Merritt, Dean, Settle.
5. Spectroscopy by G. R. Chatwal and S. K. Anand
6. Spectroscopy by Pavia, Lampman, Kriz, Vyvyan
7. Organic Spectroscopy (2nd edition) by Jag Mohan
8. Organic Spectroscopy (3rd edition) by William Kemp
9. Instrumental Methods of Chemical Analysis by H. Kaur

B. Sc. Part III (NEP 1. 0) SEMESTER-V

Paper No. DSE-E7 Chemistry Paper No. XI (Physical Chemistry)

(Theory Credits: 02, 30 hours.)

Expected learning Outcomes:

Name of the Topics	Expected Learning Outcome
1.Elementary quantum mechanics	Learning and understanding quantum Chemistry, Heisenberg's uncertainty principle, concept of energy operators (Hamiltonian), learning of Schrodinger wave equation. Physical interpretation of the ψ and ψ^2 . Particle in a one dimensional box
2. Spectroscopy	Knowledge about spectroscopy, Electromagnetic spectrum, Energy level diagram, Study of rotational spectra of diatomic molecules: Rigid rotor model, Microwave oven, vibrational spectra of diatomic molecules, simple Harmonic oscillator model, Raman spectra: Concept of polarizability, pure rotational and pure Vibrational Raman spectra of diatomic molecules, related knowledge will be gained by the students.
3. Photochemistry	Learning and understanding photochemical laws, reactions and various photochemical phenomena.
4. Solution	Learning the various types of solutions, relations vapour pressure, temperature relations.
5. Electromotive force	Learning and understanding the knowledge of emf measurements, types of electrodes, different types of cells, various applications of emf measurements.

Unit 1. Elementary quantum mechanics

[5]

- 1.1 Introduction.
- 1.2 Drawbacks of classical mechanics, Black body radiation, Photoelectric effect, Compton effect,
Duel nature of matter and energy: De Broglie hypothesis.
- 1.3 The Heisenberg's uncertainty principle.
- 1.4 Concept of energy operators (Hamiltonian).
- 1.5 Derivation of Schrodinger wave equation, well behaved function.
- 1.6 Physical interpretation of the ψ and ψ^2 .
- 1.7 Particle in a one dimensional box.
- 1.8 Numerical problems.

Unit 2. Spectroscopy

[7]

- 2.1 Introduction.
- 2.2 Electromagnetic radiation.
- 2.3 Interaction of radiation with matter, Electromagnetic spectrum, Energy level diagram.
- 2.4 Rotational spectra of diatomic molecules: Rigid rotor model, moment of inertia, energy levels of rigid rotor, selection rules, Intensity of spectral lines, determination of bond length, isotope effect, Microwave oven
- 2.5 Vibrational spectra of diatomic molecules: Simple Harmonic oscillator model,
Vibrational energies of diatomic molecules, Determination of force constant, overtones.
- 2.6 Raman spectra: Concept of polarizability, pure rotational and pure Vibrational Raman spectra of diatomic molecules, selection rules.
- 2.7 Comparative study of IR and Raman spectra, rule of mutual exclusion- CO₂ molecule.
- 2.8 Numerical problems.

Unit 3. Photochemistry

[5]

- 3.1 Introduction, Difference between thermal and photochemical processes.
- 3.2 Laws of photochemistry: i) Grotthus - Draper law ii) Lambert law iii) Lambert – Beer's law (with derivation) iv) Stark-Einstein law.
- 3.3 Quantum yield, Reasons for high and low quantum yield.
- 3.4 Factors affecting Quantum yield.
- 3.5 Photosensitized reactions – Dissociation of H₂, Photosynthesis.
- 3.6 Photodimerisation of anthracene, decomposition of HI and HBr.
- 3.7 Jablonski diagram depicting various processes occurring in the excited state:
Qualitative description of fluorescence and phosphorescence.

3.8 Chemiluminescence, Electroluminescence and Bioluminescence.

3.9 Numerical problems.

Unit 4. Solutions

[5]

4.1 Introduction.

4.2 Ideal solutions, Raoult's law, Vapour pressure of ideal and non ideal solutions of miscible liquids.

4.3 Composition of liquid and vapour, vapour pressure and boiling point diagrams of miscible liquids. Distillation of miscible liquid pairs.

Type I : Systems with intermediate total vapour pressure (i.e. System in which b.p. increases regularly – Zeotropic).

Type II : Systems with a maximum in the total vapour pressure (i.e. System with a b.p. minimum – Azeotropic).

Type III : Systems with a minimum in the total vapour pressure (i.e. System with a b.p. Maximum – Azeotropic).

4.4 Solubility of partially miscible liquids.

- i. Maximum solution temperature type: Phenol – water system.
- ii. Minimum solution temperature type: Triethyl amine – water system.
- iii. Maximum and minimum solution temperature type: Nicotine – water system.

Distillation of partially miscible liquid pairs.

4.5 Vapour pressure and distillation of immiscible liquids, steam distillation.

Unit 5. Electromotive force

[8]

(Convention: Reduction potentials to be used)

5.1 Introduction

5.2 Thermodynamics of electrode potentials, Nernst equation for electrode and cell potentials in terms of activities.

5.3 E.M.F. series.

5.4 Types of electrodes: Description in terms of construction, representation, half-cell reaction and emf equation for

- i) Metal – metal ion electrode.
- ii) Amalgam electrode.
- iii) Metal – insoluble salt electrode.
- iv) Gas – electrode.
- v) Oxidation – Reduction electrode.

5.5 Reversible and Irreversible cells.

- i. Chemical cells without transference.
- ii. Concentration cells with and without transference.
- iii. Liquid – Liquid junction potential: Origin, elimination and determination.

5.6 Equilibrium constant from cell emf, Determination of the thermodynamic parameters such as ΔG , ΔH and ΔS .

5.7 Applications of emf measurements:

- i. Determination of pH of solution using Hydrogen electrode.
- ii. Solubility and solubility product of sparingly soluble salts (based on concentration cells).

5.8 Numerical problems.

Reference Books:

1. Physical Chemistry by G. M. Barrow, International student Edition, Mc Graw Hill.
2. University General Chemistry by C.N.R. Rao, Macmillan.
3. Physical Chemistry by, R. A. Alberty, Wiley Eastern Ltd.
4. The Elements of Physical Chemistry by P. W. Atkins, Oxford.
5. Principles of Physical Chemistry by S. H. Maron, C. H. Prutton, 4th Edition.
6. Nuclear and Radiochemistry by Friedlander, Kennedy and Miller, John Wiley and Sons.
Wiley International edition.
7. Essentials of Nuclear Chemistry by H. J. Arnikar, 4th edition. Wiley Eastern.
8. Principles of Physical Chemistry by Puri, Sharma, Pathania, Shobhanlal Naginchand and Company, Jalandar.
9. Instrumental methods of chemical analysis by Chatwal and Anand, 5th Edition, Himalaya Publication.
10. Fundamentals of molecular spectroscopy by C. N. Banwell – Tata Mc Graw-Hill.
11. Quantum Chemistry including molecular spectroscopy by B. K. Sen, Tata Mc Graw -Hill.
12. Text Book of Physical Chemistry by S. Glasstone, Macmillan India Ltd.
13. Elements of Physical Chemistry by D. Lewis and S. Glasstone (Macmillan).
14. Principles of Physical Chemistry by Maron and Lando (Amerind).
15. Electrochemistry by S. Glasstone.
16. Physical Chemistry by W. J. Moore.

17. Basic Chemical Thermodynamics by V. V. Rao (Macmillan).
18. Essential of Physical Chemistry, Bahl and Tuli (S. Chand).
19. Text Book of Physical Chemistry, Soni and Dharmarha.
20. Advanced Physical Chemistry Gurdeep Raj GOEL Publishing House, 36th Edition

B. Sc. Part III (NEP 1. 0) SEMESTER-V

Paper No. DSE-E8 Chemistry Paper No. XII (Analytical Chemistry)

(Theory Credits: 02, 30 hours.)

EXPECTED LEARNING OUTCOMES:

Name of the topic	Expected Learning Outcome
1.Potentiometric Titrations	Understanding theory and applications of potentiometric titrations.
2.Colorimetry and Spectrophotometry	Understanding, working and applications of optical methods as an analytical tool.
3.Sugar Industry	Learning and understanding the whole process of manufacture of sugar and byproducts of sugar industry.
4.Manufacture of Industrial Heavy Chemicals	Learning and understanding of physico-chemical principles of production of ammonia, sulfuric acid, nitric acid and sodium carbonate along with its manufacturing plant.
5. Gas Chromatography and Quality Control	Understanding the basics of Gas Chromatography, Quality control practices in analytical industries / laboratories.

- 1. Potentiometric titrations** [7]
- 1.1 Introduction.
 - 1.2 Determination of pH.
 - 1.3 Study of Quinhydrone and Glass electrodes and their use in determination of pH.
 - 1.4 Potentiometric titrations: Classical and analytical methods for locating end points.
 - 1.5 Acids- Bases titration with suitable example.
 - 1.6 Redox titration with suitable example.
 - 1.7 Precipitation titration with suitable example.
 - 1.8 Basic circuit of direct reading potentiometer.
 - 1.9 Advantages of potentiometric titrations
- 2. Colorimetry and Spectrophotometry** [5]
- 2.1 Theory of colorimetry and spectrophotometry.
 - 2.2 Lambert Beer's law, deviation from Beer's law.
 - 2.3 Terms used in colorimetry and spectrophotometry.
 - 2.4 Classification of methods of 'colour' measurement or comparison.
 - 2.5 Photoelectric colorimeter method—Single beam photo-electric colorimeter.
 - 2.6 Spectrophotometer method—Single beam direct reading spectrophotometer.
 - 2.7 Determination of unknown concentration by using concentration-absorbance plot.
 - 2.8 Applications of colorimetry and spectrophotometry.
- 3. Sugar Industry** [5]
- 3.1 Introduction.
 - 3.2 Manufacture of cane sugar in India: Extraction of juice, Clarification, Concentration, crystallization, centrifugation and other details of industrial process.
 - 3.3 Byproducts of sugar industry.
 - 3.4 Manufacture of Ethyl Alcohol from Molasses: by Fermentation.
- 4. Manufacture of industrial heavy chemicals** [6]
- 4.1 Introduction
 - 4.2 Manufacture of Ammonia (NH_3)
 - 4.2.A Physico-chemical principles.
 - 4.2.B Manufacture by Haber's process.

4.3 Manufacture of Sulphuric acid (H_2SO_4)

4.3.A Physico-chemical principles.

4.3.B Manufacture by Contact process.

4.4 Manufacture of Nitric acid (HNO_3)

4.4.A Physico-chemical principles.

4.4.B Manufacture by Ostwald's process (Ammonia oxidation process).

5. A. Gas Chromatography

[7]

5.1 Introduction to Chromatography

5.2 Principle of GC

5.4 Instrumentation of GC

5.4 Applications of GC

5.5 Advantages and disadvantages

B. Quality control: Concepts in Quality control

- i. Introduction and Concept of quality.
- ii. Quality control.
- iii. Quality assurance.
- iv. ISO series.
- v. Good laboratory practices.

C. IPR : Industrial design and patent

References

1. Text Book of Quantitative inorganic analysis – A.I.Vogel.
2. Instrumental methods of chemical analysis –Willard, Merit & Dean.
3. Instrumentals methods of chemical analysis – Chatwal & Anand.
4. Vogel's textbook of qualitative inorganic analysis – Bassett, Denny etc.
5. Fundamentals of analytical chemistry – Skoog and West.
6. Basic concepts of analytical chemistry – S.M. Khopkar.
7. Text book of qualitative chemical analysis – Vogel.
8. Handbook of quality assurance for the analytical chemistry laboratory
– James P.Dux, Van Nostrand Reinhold, New York 1986.
9. Instrumental methods of chemical analysis – H.Kaur.
10. A text book of Quantitative chemical analysis Vogel's by J.Mendham, R. C. Denney.

11. Quantitative Chemical Analysis – Daniel C. Harris.
12. Applying ISO 9000 Quality management system, International trade centre publishing genera, Indian edition printed by D. L. Shaha Trust.
13. Industrial Chemistry-B.K. Sharma
14. Chemical process industries – Shrieve & Brink
15. Industrial chemistry – Kent
16. Industrial chemistry – Rogers
17. Industrial chemistry – R. K. Das

B. Sc. Part III (NEP 1. 0) SEMESTER-VI

Paper No. DSE-F5 Chemistry Paper No. XIII (Inorganic Chemistry)

(Theory Credits: 02, 30 hours.)

EXPECTED LEARNING OUTCOME:

Name of the topic	Expected Learning Outcome
1. Coordination Chemistry	The topic focused on the mechanism of the reactions involved in inorganic complexes of transition metals. The students can understand the thermodynamic and kinetic aspects of Metal complexes.
2. Nuclear Chemistry	The generation of nuclear power with the help of nuclear reactions is highlighted. Role of Radio isotopes in medicinal, industrial and Archaeology fields is explained.
3. Chemistry of f-block Elements	The characteristics properties and separation of lanthanides and Actinides are discussed. Synthesis and IUPAC Nomenclature of trans -uranic elements (TU) explained.
4. Iron and Steel	The techniques involve in ore dressing and extraction of cast iron from its ore are discussed.
5. Bio-inorganic Chemistry	Role of various metals and non metals in our Health are discussed.

Unit- 1. Coordination Chemistry

[9]

A. Inorganic Reaction mechanism

- 1.1 Introduction.
- 1.2 Classification of Mechanism: Association, dissociation, interchange and the rate determining steps.
- 1.3 S_N^1 and S_N^2 reactions for inert and labile complexes.
- 1.4 Mechanism of substitution in cobalt(III) octahedral complexes.
- 1.5 Trans effect and its theories.
- 1.6 Applications of trans effect in synthesis of Pt(II) complexes.

B. Thermodynamic and Kinetic aspects of metal complexes.

- 1.7 Introduction.
- 1.8 Thermodynamic stability.
- 1.9 Kinetic Stability.
- 1.10 Relation between thermodynamic and kinetic stability.
- 1.11 Stepwise stability constant.
- 1.12 Factors affecting the stability of complexes.
- 1.13 Determination of Stability constant by Job variation, Mole ratio and Slope ratio method.

Unit-2. Nuclear Chemistry

[4]

- 2.1 Nuclear reactions and energetic of nuclear reactions.
- 2.2 Types of nuclear reactions
 - 2.2.1 Artificial transmutation.
 - 2.2.2 Artificial radioactivity.
 - 2.2.3 Nuclear fission and its application in heavy water nuclear reactor.
 - 2.2.4 Nuclear fusion.
- 2.3 Use of Thorium, Uranium and Plutonium in atomic energy
- 2.4 Applications of radio-isotopes as tracers.
 - 2.4.1 Chemical investigation—Esterification.
 - 2.4.2 Structural determination—Phosphorus pentachloride.
 - 2.4.3 Analytical Chemistry—Isotopic dilution method for determination of volume of blood.
 - 2.4.4 Age determination—Dating by C^{14} .

Unit-3. Chemistry of f-Block Elements

[7]

A] Lanthanides

- 3.1 Introduction.
- 3.2 Occurrence.
- 3.3 Electronic Configuration.
- 3.4 Oxidation State.
- 3.5 Lanthanide contraction.
- 3.6 Separation of Lanthanides by Ion exchange method.

B] Actinides

- 3.7 Position in periodic table.
- 3.8 Electronic configuration.
- 3.9 General methods of preparation of transuranic elements.
 - 3.9.1 Neutron capture—followed by β -decay.
 - 3.9.2 Accelerated projectile bombardment.
 - 3.9.3 Heavy ion bombardment.
- 3.10 IUPAC nomenclature of the super heavy elements with atomic number(Z) greater than 100.

Unit-4. Iron and Steel.

[6]

- 4.1 Occurrence and ores of iron.
- 4.2 Definition of the Terms-Ore, Mineral, Slag, Flux, Gangue, Matrix, Calcinations, Reduction, Roasting, Smelting and Leaching.
- 4.3 Extraction of iron by Blast furnace.
- 4.4 Steel: Definition and types.
- 4.5 Conversion of cast iron into steelby
 - 4.5.1 Bessemer process.
 - 4.5.2 L. D. process.
- 4.6 Heat treatment on steel.

Unit-5. Bio-inorganic Chemistry.

[4]

- 5.1 Introduction.
- 5.2 Essential and trace elements in biological process.
- 5.3 Metalloporphyrins with special reference to hemoglobin and myoglobin.
- 5.4 Biological role of alkali and alkaline earth metal ions with special reference to Na^+ , K^+ and Ca^{2+}

Reference Books: (Use recent editions)

1. Concise Inorganic Chemistry (ELBS, 5th Edition)—J. D. Lee.
2. Inorganic Chemistry (ELBS, 3rd Edition) D. F. Shriver, P. W. Atkins, C. H. Langford, Oxford University Press, 2nd Edition.
3. Basic Inorganic Chemistry: Cotton and Wilkinson.
4. Advanced Inorganic Chemistry (4th Edn.) Cotton and Wilkinson.
5. Concepts and Models of Inorganic Chemistry: Douglas and McDaniel. 3rd Edition. John Wiley publication.
6. Structural principles in inorganic compounds. W. E. Addison.
7. Theoretical principles of Inorganic Chemistry—G. S. Manku.
8. Theoretical Inorganic Chemistry by Day and Selbina.
9. Co-ordination compounds. SFA Kettle.
10. Essentials of Nuclear Chemistry by H. J. Arnikar.
11. Nuclear Chemistry by M. N. Sastri
12. Organometallic Chemistry by R. C. Mahrotra A. Sing, Wiley Eastern Ltd. New Delhi.
13. Inorganic Chemistry by A. G. Sharpe, Addison—Wesley Longman—Inc.
14. Principles of Inorganic Chemistry by Puri, Sharma and Kalia, Vallabh Publication. Pitampur Delhi.
15. Text book of Inorganic Chemistry by K. N. Upadhyaya Vikas Publishing House—New Delhi.
16. Inorganic Chemistry 3rd edn G. L. Miessler and D. A. Tarr, Pearson publication
17. Co-ordination compounds by Baselo and Pearson.
18. UGC Inorganic chemistry by H. C. Khera, Pragati prakashan
19. UGC Advance Inorganic Chemistry by Agarwal and Keemtilal, Pragati Prakashan

B.Sc. Part III (NEP 1. 0) SEMESTER-VI

Paper No. DSE-F6 Chemistry Paper No. XIV (Organic Chemistry)

(Theory Credits: 02, 30 hours.)

Expected learning Outcomes:

Name of the Topic	Expected Learning Outcome
1. Reagents and Reactions in Organic Synthesis	Knowledge of reagents used in organic transformations and various reactions used in organic synthesis.
2. Retrosynthesis	Knowing basic terms used in retrosynthetic analysis, retrosynthesis of some organic compounds.
3. Electrophilic addition to $>C=C<$ and $-C\equiv C-$ bond	Student will learn addition reaction across $>C=C<$ bond w.r.t. hydrohalogenation, hydration hydroxylation, ozonolysis and addition of halogen, halogen acid, hydrogen, water, etc. across $-C\equiv C-$ bond.
4. Natural Products	Knowledge of terpenoids and alkaloids w.r.t. occurrence, isolation, characteristics and classification. Analytical and synthetic evidences of Citral and Nicotine.
5. Pharmaceuticals	Understanding classification of drugs, Qualities of ideal drug. Synthesis and uses of some representative drugs and Drug action of sulpha drugs.

Unit 1. Reagents and Reactions in Organic Synthesis

[8]

I] Reagents

Preparation and Applications of following reagents.

A) Oxidizing reagents

a) Osmium tetroxide. b) Selenium dioxide (SeO_2).

B) Reducing reagents

a) Lithium Aluminum Hydride LiAlH_4 . b) Raney Nickel.

II] Name Reactions

Statement, General Reaction, Mechanism and Synthetic applications of following reactions....

1. Diels -Alder reaction.
2. Meerwein –Pondorff-Verley reduction.
3. Hofmann rearrangement.
4. Wittig reaction.
5. Wagner- Meerwein rearrangement.
6. Baeyer Villiger oxidation.

Unit 2. Retrosynthesis

[5]

- 2.1 Introduction.
- 2.2 Recapitulation of basics of reaction mechanism and reagents.
- 2.3 Terms used- Target molecule (TM), Disconnection, Synthons, Synthetic equivalence, Functional group interconversion (FGI), one group disconnection (w. r. t. suitable examples).
- 2.4 Retrosynthetic analysis and synthesis of target molecules: Cinnamaldehyde, Cyclohexene, para methoxy acetophenone, Methyl-3-phenyl propionate, α,α -dimethyl benzyl alcohol, Paracetamol.

Unit 3. Electrophilic addition to $>\text{C}=\text{C}<$ and $-\text{C}\equiv\text{C}-$ bonds

[6]

A. Addition to Carbon-Carbon double ($>\text{C}=\text{C}<$) bond:

- 3.1 Introduction.
- 3.2 Examples of addition reactions.
- 3.3 Mechanism of electrophilic addition to $>\text{C}=\text{C}<$ bond, orientation & reactivity,
 - i. Hydrohalogenation.
 - ii. Anti-Markovnikoff's addition (peroxide effect).
 - iii. Rearrangements (support for formation of carbocation).

- iv. Addition of halogens.
- v. Addition of water.
- vi. Addition of hypohalous acids (HO-X).
- vii. Hydroxylation (formation of 1,2-diols).
- viii. Hydroboration-oxidation (formation of alcohol).
- ix. Hydrogenation (formation of alkane).
- x. Ozonolysis (formation of aldehydes & ketones).

B. Addition to Carbon-Carbon triple ($\text{C}\equiv\text{C}$) bond:

3.4 Introduction.

3.5 Examples of addition reactions.

3.6 Mechanism of electrophilic addition to $\text{C}\equiv\text{C}$ -bond.

- i. Addition of halogens.
- ii. Addition of halogen acids.
- iii. Addition of hydrogen.
- iv. Addition of water.
- v. Formation of metal acetylides.

Unit 4. Natural Products

[6]

A] Terpenoids:

4.1 Introduction, Occurrence, Isolation, General Characteristic, Classification.

4.2 General Methods for structure determinations.

4.3 Isoprene rule.

4.4 Analytical evidences and synthesis of Citral.

B] Alkaloids:

4.5 Introduction, Occurrence, Isolation, Classification, Properties.

4.6 General Methods for structure determination.

4.7 Analytical evidences and synthesis of Nicotine.

Unit 5. Pharmaceuticals

[5]

5.1 Introduction.

5.2 Classification of drugs.

5.3 Qualities of ideal drug.

5.4 Synthesis and uses of ethambutal, phenobarbitone, isoniazide, benzocaine, Chloramphenicol and paludrine.

5.5 Drug action of sulpha drugs.

Reference Books:

1. Advanced Organic Chemistry : Reactions, Mechanisms and structure by – Jerry March.
2. Reagents for Organic Synthesis by Louis F. Fieser , Mary Fieser -1967.
3. A Text book of Practical Organic Chemistry including Qualitative Organic Analysis by A. I.Vogel.
4. Mechanism and Structure in Organic Chemistry. April,1963 By Edwin S.Gould.
5. A text book of Organic Chemistry by Arun Bahl, B.S.Bhal Eighteenth Revised edition 2006.
6. A guidebook to mechanism in Organic Chemistry sixth Edition by Peter Syke.
7. Organic Synthesis: The Disconnection Approach by Stuart Warren.
8. Organic Synthesis Through Disconnection Approach by P. S. Kalsi
9. Fundamentals of Organic Synthesis the Retrosynthetic Analysis by Ratan Kumar Kar
10. Organic Reactions and Their Mechanisms P. S. Kalsi 3rd Revised edition.
11. Advanced organic Chemistry by B.S. Bahl & Arun Bhal (Reprint in 1997)
12. Organic Chemistry by Morrison and Boyd 6th edition.
13. Organic Chemistry Vol II Stereochemistry and the Chemistry of Natural Products (5th ed) by I. L.Finar.
14. Organic Chemistry Natural Products Vol I, by O. P.Agrawal
15. Industrial Chemistry-B.K. Sharma, Goyal publishing house,Mirut
16. Shreeves chemical process industries 5th Edition, G.T. Oustin, McGrawHill
17. Riegel`s hand book of Industrial chemistry, 9th Edition, Jems A.Kent
18. Industrial chemistry –R.K. Das, 2nd Edition,1976.
19. Synthetic drugs by M.S.Yadav,Campus book international.

B.Sc. Part III (NEP 1. 0) SEMESTER-VI

Paper No. DSE-F7 Chemistry Paper No. XV Physical Chemistry)

(Theory Credits: 02, 30 hours.)

EXPECTED LEARNING OUTCOMES:

Name of the Topics	Expected Learning Outcome
1. Phase equilibria	Learning and understanding of phase rule, learning of One component, Two component and Three component systems phase diagrams with suitable examples.
2. Thermodynamics	Knowledge about basic concept of Thermodynamics, free energy, Gibbs-Helmholtz equation and its applications, problem related with it.
3. Solid state chemistry	Learning and understanding Space lattice, lattice sites, Lattice planes, Unit cell. Laws of crystallography, Weiss indices and Miller indices, Cubic lattices and types of cubic lattice, planes or faces of a simple cubic system, Diffraction of Xrays, Derivation of Bragg's equation. Determination of crystal structure by Bragg's method. crystal structure of NaCl and KCl on the basis of Bragg's equation..
4. Chemical kinetics	Learning of kinetics, Simultaneous reactions such as i)opposing reaction ii)side reaction iii)consecutive reactions: iv) chain reaction v) explosive reaction
5. Colloidal State	Learning and understanding the knowledge of Colloidal State, understanding of colloidal system, different types of colloidal system, preparation, properties, stability of different colloidal system, General applications of colloids.

Unit 1. Phase Equilibria

[6]

- 1.1 Introduction.
- 1.2 Gibbs phase rule: Phase rule equation and explanation of terms involved in the equation.
- 1.3 Phase diagram, true and metastable equilibria.
- 1.4 One component systems:
 - i. Water system.
 - ii. Sulphur system with explanation for polymorphism.
- 1.5 Two component systems:
 - i. Eutectic system: (Ag – Pb system); Desilverisation of lead.
 - ii. Freezing mixture: (KI –H₂O system).
 - iii. Formation of compound with congruent melting point (FeCl₃ – H₂O).
- 1.6 Three component solid-liquid system:
 - i. Development of triangular phase diagram: (Acetic acid – Chloroform –water system).

Unit 2. Thermodynamics

[7]

- 2.1 Introduction.
- 2.2 Free energy: Gibbs function (G) and Helmholtz function (A), Criteria for thermodynamic equilibrium and spontaneity.
- 2.3 Relation between ΔG and ΔH : Gibbs-Helmholtz equation.
- 2.4 Phase equilibria : Clapeyron – Clausius equation and its applications.
- 2.5 Thermodynamic derivation of law of mass action, Van't – Hoff isotherm and isochore.
- 2.6 Fugacity and activity concepts.
- 2.7 Partial molar quantities, Partial molar volume, Concept of chemical potential, Gibbs-Duhem equation.
- 2.8 Numerical problems.

Unit 3. The Solid State

[7]

- 3.1 Introduction: Space lattice, lattice sites, lattice planes, unit cell.
- 3.2 Laws of crystallography:
 - i. Law of constancy of interfacial angles
 - ii. Law of rational indices
 - iii. Law of crystal symmetry.
- 3.3 Weiss indices and Miller indices.
- 3.4 Cubic lattice and types of cubic lattice, planes or faces of a simple cubic system,

spacing of lattice planes.

3.5 Diffraction of X-rays, Derivation of Bragg's equation.

3.6 Determination of crystal structure by Bragg's method.

3.7 Determination of crystal structure of NaCl and KCl on the basis of Bragg's equation.

3.8 Numerical problems.

Unit 4. Chemical Kinetics

[5]

4.1 Introduction.

4.2 Simultaneous reactions such as

- i. Opposing reaction: (Derivation of rate equation for first order opposed by first order expected).
- ii. Side reaction.
- iii. Consecutive reactions.
- iv. Chain reaction.
- v. Explosive reaction (Derivation of rate equation and Numerical problems are not expected).

Unit 5. Colloidal State

[5]

5.1 Introduction: Definition of Colloids, Colloidal state of matter, True solution, colloidal solution and Suspension, Dispersed phase and Dispersion medium,

5.2 Types of Colloidal systems.

5.3 Solids in liquids (Sols): preparation, purification, properties and stability

5.4 Liquids in Liquids (Emulsions): Types and Preparations.

5.5 Liquids in Solids (Gels): Types and Preparations.

5.6 General Applications of Colloids.

Reference Books:

1. Physical Chemistry by G. M. Barrow, International student Edition, Mc Graw Hill.
2. University General Chemistry by C.N.R. Rao, Macmillan.
3. Physical Chemistry by, R. A. Alberty, Wiley Eastern Ltd.
4. The Elements of Physical Chemistry by P. W. Atkins, Oxford.
5. Principles of Physical Chemistry by S. H. Maron, C. H. Prutton, 4th Edition.
6. Nuclear and Radiochemistry by Friedlander, Kennedy and Miller, John Wiley and Sons.
Wiley International edition.
7. Essentials of Nuclear Chemistry by H. J. Arnikar, 4th edition. Wiley Eastern.
8. Principles of Physical Chemistry by Puri, Sharma, Pathania, Shobhanlal Naginchand and Company, Jalandar.

9. Instrumental methods of chemical analysis by Chatwal and Anand, 5th Edition, Himalaya Publication.
10. Fundamentals of molecular spectroscopy by C. N. Banwell – Tata Mc Graw-Hill.
11. Quantum Chemistry including molecular spectroscopy by B. K. Sen, Tata Mc Graw -Hill.
12. Text Book of Physical Chemistry by S. Glasstone, Macmillan India Ltd.
13. Elements of Physical Chemistry by D. Lewis and S. Glasstone (Macmillan).
14. Principles of Physical Chemistry by Maron and Lando (Amerind).
15. Electrochemistry by S. Glasstone.
16. Physical Chemistry by W. J. Moore.
17. Basic Chemical Thermodynamics by V. V. Rao (Macmillan).
18. Essential of Physical Chemistry, Bahl and Tuli (S. Chand).
19. Text Book of Physical Chemistry, Soni and Dharmarha.
20. Advanced Physical Chemistry Gurdeep Raj GOEL Publishing House, 36th Edition.

B.Sc. Part III (NEP 1. 0) SEMESTER-VI

Paper No. DSE-F8 Chemistry Paper No. XVI (Analytical Chemistry)

(Theory Credits: 02, 30 hours.)

Expected learning Outcomes:

Name of the topic	Expected Learning Outcome
1.Soap and Detergents	Knowledge about the chemical nature and cleansing action of soap
2.Synthetic polymers	Understanding and learning the classification, synthesis and applications of various polymers.
3.Flame Photometry	Knowledge of instrumental analysis of alkali and alkaline earth elements.
4. Soil and Fertilizer Analysis	Knowledge of analysis of soil essential parameters and nutrients, Understanding the fertilizers used in regular farming
5.Nanotechnology	Understanding and learning of nanotechnology including classification, optical properties, synthesis routes, characterization techniques and applications of nano-materials.

1. Soap and Detergents

[6]

- 1.1 Introduction,
- 1.2 Soaps - Raw materials,
- 1.3 Types of soaps,
- 1.4 Cleansing action of soap,
- 1.5 Manufacture of soap - Boiled or Hot Process,
- 1.6 Detergents - Raw Materials,
- 1.7 Types of Detergents: Anionic, cationic and amphoteric,
- 1.8 Preparation of Teepol and Deriphat,
- 1.9 Comparisons between soaps and detergents.

2. Synthetic polymers

[6]

- 2.1 Introduction.
- 2.2 Classification.
 - 2.2.A Based on origin.
 - 2.2.B Based on composition-organic, inorganic polymers.
 - 2.2.C Based on method of preparation.
 - 2.2.D Based on general physical properties.
 - 2.2.E Based on structure.
- 2.3 Addition Polymerization: Free radical addition and ionic addition polymerization.
- 2.4 Zigler Nata polymerization.
- 2.5 Methods of preparation and applications of some organic polymers: Polyethylene, polystyrene, polyvinyl chloride, Phenol-formaldehyde resin.
- 2.6 Conducting organic polymers: Synthesis and properties of Polyaniline, polypyrrol.
- 2.7 Applications of conducting organic polymers.

3. Flame Photometry

[5]

- 3.1 Introduction.
- 3.2 General principles of flame photometry.
- 3.3 Instrumentation: Block diagram, Burners (Premix and Lundergraph burners), mirror, slits, filters, detector (Photomultiplier tube).
- 3.4 Experimental procedure of analysis (Standard addition and internal standard).
- 3.5 Interferences and Factors that influence the intensity of emitted radiation

in a flame photometer.

3.6 Applications of flame photometry in real sample analysis.

3.7 Limitations of flame photometry.

4. Fertilizer and Soil Analysis

[8]

4.1 Introduction

4.2 Need of Fertilizers

4.3 Classification of fertilizers

4.4 Source of Fertilizers

A) Natural Organic Fertilizer

B) Natural Inorganic Fertilizer

C) Artificial Fertilizer

(Fertilizer Analysis is similar to soil analysis)

4.5 Components of soil

A) Macro and micro plant nutrients

B) Functions of nutrients

4.6 Analysis of Soil

A) Soil moisture by gravimetric method

B) Soil pH

C) Total Nitrogen by kjeldahl method

D) Total phosphorous

E) Potassium by flame photometry

F) Sodium by flame photometry

B Nanotechnology

[5]

5.1 Introduction of nanotechnology, history, Classification of nanoparticles based on size.

5.2 Synthetic Routes of nanomaterials: Top-down and bottom-up approaches.

5.3 Synthesis methods: Sol-gel, precipitation, chemical reduction, chemical vapor deposition, hydrothermal, electrodeposition.

5.4 Characterization of nanomaterials: X-Ray diffractometer

5.5 Scanning Electron Microscope, Transmission electron microscope.

5.6 Applications of nanotechnology.

References:

1. Industrial Chemistry-B.K. Sharma
2. Chemical process industries – Shrieve & Brink
3. Industrial chemistry – Kent
4. Industrial chemistry – Rogers
5. Industrial chemistry – R. K. Das
6. Nanotechnology: Principles and Practices – Sulbha Kulkarni
7. The Petroleum chemicals industry by R. F. Goldstine, e &Fn London
8. Fundamentals of petroleum chemical technology - P Below.
9. Petro Chemicals Volume 1 and 2 ; A Chauvel and Lefevrev ; Gulf Publishing company.
10. Text Book of Quantitative inorganic analysis – A.I.Vogel.
11. Instrumental methods of chemical analysis –Willard, Merit & Dean.
12. Instrumentals methods of chemical analysis – Chatwal & Anand.
13. Vogel’s textbook of qualitative inorganic analysis – Bassett, Denny etc.
14. Basic concepts of analytical chemistry – S.M. Khopkar.
15. Text book of qualitative chemical analysis – Vogel.
16. Instrumental methods of chemical analysis – H.Kaur.
17. Quantitative Chemical Analysis – Daniel C. Harris.

Laboratory Course (Practicals)

N. B. (i) Use of Digital/Analytical/Chainometric/Single pan balance is allowed.

(ii) Use of Scientific calculator is allowed.

(iii) Use of Chart/Text book/Hand book of practical is allowed.

(iv) There will be a project having weightage of 15 marks.

Project should be in the following areas but focused on applications of Chemistry.

a) Society oriented

b) Daily use

c) Industry based

d) Analysis based

The project will be assessed by all the three examiners with equal weightage at the time of practical examinations. The project may be completed individually or by a group of students not exceeding number three. One copy of the project should be submitted at the time of examination. After assessment this copy will remain in the department.

INORGANIC CHEMISTRY

I) Gravimetric Estimations (G).

N. B. Any two experiments from G1 to G3 and any two experiment from G4 & G6.

G1. Gravimetric estimation of iron as ferric oxide (Fe_2O_3) from the given solution containing ferrous ammonium sulphate, copper sulphate and free sulphuric acid.

G2. Gravimetric estimation of zinc as zinc pyrophosphate from the given solution containing zinc sulphate, ferrous ammonium sulphate and free sulphuric acid.

G3. Gravimetric estimation of barium as barium sulphate(BaSO_4) from the given solution containing barium chloride, ferric chloride and free hydrochloric acid.

G4. Gravimetric estimation of barium as barium chromate(BaCrO_4) from the given solution containing barium chloride, ferric chloride and free hydrochloric acid.

G5. Gravimetric estimation of nickel as bis (dimethylglyoximate) nickel (II) from the given solution containing nickel sulphate, ferrous ammonium sulphate and free Sulphuric acid.

G6. Gravimetric estimation of aluminium as aluminium oxinate potassium tris (8-hydroxy quinolato) aluminium (III) from the given solution containing potash alum ,copper sulphate and free sulphuric acid.

[For the gravimetric experiments, stock solution should be given in the range of 10 to 15 cm³ and asked to dilute to 100 cm³ (or the stock solution should be given in the range of 20 to 30 cm³ and asked to dilute to 250 cm³). Use 50 cm³ of this diluted solution for estimation.]

II. Inorganic Preparations (P).

N. B. At least six preparations from the following with percentage yield:

- P1. Preparation of potassium trioxalato aluminate (III).
- P2. Preparation of Tetra ammine copper (II) chloride.
- P3. Preparation of tris(thiourea) copper (I) sulphate.
- P4. Preparation of potassium trioxalato ferrate (III).
- P5. Preparation of chloropenta-ammine cobalt (III) chloride.
- P6. Preparation of ammonium diamminetetraethiocyanato chromate (III) (Reineck's salt).
- P7. Preparation of Potassium hexa nitro cobaltate (III).
- P8. Preparation of ammonium trioxalato chromate (III).
- P9. Preparation of hexathiourea plumbus (II) nitrate.

A) Percentage Purity

N. B. : Any two from the following.

- V1. Determination of percentage purity of ferrous ammonium sulphate.
- V2. Determination of percentage purity of tetrammine copper (II) sulphate.
- V3. Determination of percentage purity of potassium (trioxalato-aluminate) (III).

B) Analysis of Commercial Sample.

N. B. Any Three from the following:

- V5. Determination of percentage of Calcium in the given sample of milk powder or lime.
- V6. Determination of amount of aluminum in the given solution of potash alum.
- V7. Determination of titrable acidity in the given sample of milk or lassi.
- V8. Determination of percentage purity of boric acid using supplied sodium hydroxide.

(Standard succinic or oxalic acid solution to be prepared to standardise the given sodium hydroxide solution.)

- V9. To determine the amount of HCl in given of commercial samples.

C) Ion exchange method.

N. B. Any two from the following.

- V10. Determination of amount of sodium present in the given solution of common salt using cation exchange resin (By Acid Base titration).
- V11. Determination of amount of magnesium in the given solution containing (Mg^{2+} and Zn^{2+}) using anion exchange resin and standard solution of EDTA.
- V12. Determination of amount of zinc in the given solution containing (Mg^{2+} and Zn^{2+}) using anion exchange resin and standard solution of EDTA.

Reference Books:

1. A text book of quantitative Inorganic Analysis - A. I. Vogel.
2. Text book of Quantitative Inorganic Analysis - Kolthoff and Sandell.
3. Experimental Inorganic Chemistry - Palmer W. G.
4. Advanced Practical Inorganic Chemistry - Adams and Raynor.
5. Manual in Dairy Chemistry - I.C.A.R. Sub-Committee on Dairy Education.
6. Chemical methods for environmental analysis - R. Ramesh and M. Anbu.

Organic Chemistry Practical's

I) Qualitative analysis

Separation of binary organic mixture and Identification of **TWO** components.
(At least 08 mixtures)

Nature of Binary Mixtures

- 1) Solid – Solid : 4 mixtures
- 2) Solid – Liquid : 2 mixtures
- 3) Liquid – Liquid : 2 mixtures

1) Solid – Solid Mixtures:

One mixture from each the following types should be given:

- i) Acid + Phenol ii) Acid + Base
- iii) Acid + Neutral iv) Phenol + Base
- v) Phenol + Neutral vi) Base + Neutral

2) Solid – Liquid Mixtures

Mixture of type Neutral + Neutral or Acid + Neutral should be given.

3) Liquid – Liquid Mixtures

Mixture of type Neutral + Neutral or Base + Neutral
should be given for liquid – liquid mixture.

Following compounds should be used for preparation of
binary organic mixtures.

- i) **Acids:** Benzoic acid, Phthalic acid, Salicylic acid, Cinnamic acid, Aspirin, Oxalic acid.
- ii) **Phenols:** α -naphthol, β -naphthol.
- iii) **Bases:** o-nitroaniline, m-nitroaniline, p-nitroaniline, aniline, o-toluidine and

N, N- dimethylaniline.

iv) **Neutrals:** Anthracene, acetanilide, m-dinitrobenzene, chloroform, carbon tetrachloride, acetone, nitrobenzene, ethyl acetate, ethyl benzoate, bromobenzene, urea and thiourea.

NB :

1. For Solid-Liquid and Liquid-Liquid mixtures avoid detection of type of mixture. Instead the weightage is given to detection of physical nature and separation of mixture.
2. Separation and qualitative analysis of the binary Mixtures should be carried out on microscale using microscale kits.

II) Quantitative analysis: Organic Estimations:(Any four)

1. Estimation of sucrose
2. Saponification value of oil.
3. To determine the amount of acid and amide present in the given mixture of acid and amide.
4. Determination of Molecular weight of monobasic/dibasic acid by volumetric method.
5. Estimation of unsaturation –to estimate the percentage purity of given olefinic compound by bromination method.

III) Organic Preparations: (Any four)

1. Multicomponent reaction - Preparation of Dihydropyrimidone.
2. Radical coupling reaction - Preparation of 1,1,2 bis-2naphthol.
3. Base catalyzed Aldol condensation- Preparation of Dibenzal propanone.
4. Diels Alder reaction- Reaction between Furan and Maleic acid
5. Benzil- Benzilic acid rearrangement reaction
6. Oxidation reaction – Preparation of Methyl phenyl sulfone.

IV) Determination of structure of organic compound from given NMR spectra.

Ethanol, 2-methyl butane, Ethyl acetate, Benzyl alcohol, Propanoic acid, Butanaldehyde, Ethyl benzoate, Benzoic acid, Propyl ether, n-pentane, Propene, Diethyl amine, 2-chloro butane, Toulene

NB: All preparations should be carried out by considering green Chemistry approach

1. Preparation of derivative should be carried out on small scale. The starting

compound should not be given more than one gram.

2. Calculation of percentage practical yield in preparation is must.
3. Recrystallization of crude product and its melting point.
4. The product should be confirmed by TLC.
5. Assign reactions with mechanism.
6. Spectral data should be provided to the students.

Reference books:

1. Practical Organic Chemistry by – A. I. Vogel.
2. Practical Organic Chemistry by – O. P. Agarwal

PHYSICAL CHEMISTRY

I. Non instrumental Experiments:

A. Any one of the following.

i) Partition Law.

To determine the partition coefficient of CH_3COOH between H_2O and CCl_4 .

ii) Viscosity.

To determine the viscosity average molecular weight of a polymer.

iii) Adsorption.

To investigate the adsorption of oxalic acid by activated charcoal and test the validity of Freundlich & Langmuir isotherms.

iv) Solubility.

To study the effect of addition of electrolyte (NaCl or KCl) on the solubility of Benzoic acid at room temperature.

B. Chemical kinetics. (Any four)

1. The study of energy of activation of first order reaction i.e. hydrolysis of methyl acetate in presence of 0.5 N HCl / $0.5 \text{ N H}_2\text{SO}_4$.
2. The study of energy of activation of second order reaction i.e. reaction between $\text{K}_2\text{S}_2\text{O}_8$ and KI (Equal concentrations).
3. The study of energy of activation of second order reaction i.e. reaction between $\text{K}_2\text{S}_2\text{O}_8$ and KI (Unequal concentrations).
4. To study the hydrolysis of methyl acetate by using its two concentrations in presence of

0.5 N HCl and hence find velocity constant of the reaction.

5. To study the effect of addition of electrolyte (KCl) on the reaction between $K_2S_2O_8$ and KI (Equal concentrations).

C. Partial molar volume.

1. To determine the partial molar volume of ethyl alcohol in a mixture of ethyl alcohol and water (Any seven mixtures be given).

II. Instrumental experiments

A. Potentiometry (Any four)

1. Titration of strong acid with strong alkali.
N.B. i) 8 to 10 ml of 1N acid solution to be given by examiner in 100 ml volumetric flask & student should dilute it to 100 ml and 10ml of this solution is taken for titration.
ii) Experiment is carried out by taking pilot run from 1 to 10 ml and then final run taking 0.2 ml reading in the range of end point.
2. Preparation of buffer solution and determination of their pH (Any five buffer solutions), Theoretical calculation of pH values by using Henderson's equation.
3. Determination of standard electrode potential of Zn/Zn^{++} , Cu/Cu^{++} , Ag/Ag^+ (Any two).
4. Estimate the amount of Cl^- , Br^- and I^- in given unknown halide mixture by titrating it against standard $AgNO_3$ solution.
5. Titration of ferrous ammonium sulphate using $K_2Cr_2O_7$ solution and to calculate redox potential of Fe^{++} , Fe^{+++} system.

B. Conductometry (Any three).

- N.B. i) 8 to 10 ml of 1N acid solution to be given by examiner in 100 ml volumetric flask & student should dilute it to 100 ml and 10ml of this solution is taken for titration.
1. Titration of a mixture of weak acid and strong acid with strong alkali
 2. To study the effect of substituent on dissociation constant of weak acid with respect to acetic acid and monochloroacetic acid (cell constant to be given).
N.B. Calculate K by using formula $K = \frac{\alpha^2 C}{1 - \alpha}$
 3. To determine the velocity constant of hydrolysis of ethyl acetate by NaOH solution by conduct metric method.
 4. To determine the normality of citric acid in lemon by titrating it against standard 0.2 N NaOH solution by conductometric method.
 5. To determine λ_∞ of strong electrolyte (NaCl or KCl) and to verify Onsager equation.

C. Refractometry. (Any One)

1. To determine the percentage composition of unknown mixture by (i) graphical method and (ii) by composition law (Densities of pure liquids A & B be given).
2. To determine the molar refractivity of methyl acetate, ethyl acetate, n-hexane and carbon tetrachloride and calculate the refraction equivalents of C, H and Cl atoms.

D. Colorimetry (Any Two).

1. To verify Lambert – Beer's law using CuSO_4 solution.
2. To estimate of Fe^{+++} ions by thiocyanate method.
3. To estimate Fe^{+++} ions using salicylic acid by colorimetric titration.
4. To determine the order of reaction for the oxidation of alcohol by potassium dichromate and potassium permanganate in acidic medium colorimetrically.

E. pH – metry (Any One).

1. To determine the dissociation constant of monobasic acid (Acetic acid).
2. To determine the dissociation constant of dibasic acid (Malonic acid).
3. To determine hydrolysis constant of aniline hydrochloride.

Reference Books:

1. Findlay's Practical Physical Chemistry (Longman)
2. Advanced Practical Physical Chemistry by J. B. Yadav, Goel publishing house.
3. Practical Physical Chemistry by B. D. Khosla, V. C. Garg (R. Chand and Co.)
4. Systematic experimental Physical Chemistry by Rajbhoj, Chandekar (Anjali Publicaiton) Aurangabad.
5. Practical Physical Chemistry: Nandkumari, Kothari and Lavande.
6. Practical Physical Chemistry by Gurtu (S. Chand).
7. Text Book of Qualitative Inorganic Analysis by A. I. Vogel (ELBS Longman).

Nature of Practical Examination

- 1) The practical examination will be of 200 marks.
- 2) The duration of practical examination will be of three days - six and half hour per day.
- 3) Questions related to the practical exercise/project report/industrial visit carried out by the student should be asked in viva.
- 4) Use of scientific calculator is allowed.
- 5) S.I. units should be used wherever possible.

- 6) Use of Chart / Hand book / Text book of practical is allowed.
- 7) A student is expected to submit a journal certified by the Head of the Department.
- 8) A student not be permitted to appear at the practical examination unless he/she produces a certified journal. If the journal is lost, the student should produce a certificate from the Head of the Department stating that he/she has satisfactory completed the practical work but his / her journal is lost. In this case he/she will be given Zero mark for journal.
- 9) Use of Digital / Analytical / Chainometric / Single pan balance is allowed.
- 10) A student should submit one copy of project at the time of examination.
Each examiner should asses the project work for Five marks and sign the same. If any student will not submit project work, he/she will be given Zero mark for the project.
- 11) The distribution of marks for practical examination will be as follows:

A) Physical Chemistry 60 marks :

- i) Non-instrumental experiment 25 marks
- ii) Instrumental experiment 25 marks
- iii) Viva 05 marks
- iv) Journal 05 marks

B) Inorganic Chemistry 65 marks:

- i) Gravimetric analysis 25 marks
- ii) Preparation 15 marks
- iii) Volumetric estimation 15 marks
- iv) Viva 05 marks
- v) Journal 05 marks

C) Organic Chemistry 60 marks:

- i) Mixture separation and identification of **ONE** compound (at the time of practical exam identification of any one compound is given as per guideline of examiner) 25 marks
- ii) Estimation/Preparation 20 marks
- iii) Derivative 05 marks
- iv) Viva 05 marks
- v) Journal 05 marks

D) Project 15 marks

Total:- 200 marks

List of Laboratory Equipments

Apparatus & Equipments

1. Digital balance with 1 mg accuracy
2. Conductometer
3. Potentiometer
4. pH Meter
5. Polarimeter
6. Colorimeter
7. Thermostat
8. Electric Oven
9. Suction Pump
10. Crucible Heater
11. IR Lamp
12. Magnetic stirrer
13. Buckner funnel
14. Water bath / Thermostat.
15. Platinum electrode
16. Glass electrode
17. Silver, Zinc, Copper electrodes
18. Conductivity cell
19. Distilled water plant.
20. Refractometer
21. Freeze
22. Deep Freeze
23. H₂S Apparatus
24. Muffle Furnace
25. Magnetic Stirrer

Glassware & Porcelain ware:

1. Burette (25/50 ml)
2. Micro burette (10 ml)
3. Pipette (5 ml, 10 ml, 25ml)
4. Graduated Pipette (1/2/5/10 ml)
5. Conical flask (100 ml, 250 ml)

6. Beakers (100 ml, 250 ml, 500 ml)
7. Volumetric flask (25ml, 50 ml, 100 ml, 250 ml)
8. Gooch Crucible / Sintered glass Crucible
9. Silica Crucible
10. Watch glass
11. Glass tubing
12. Glass Funnel (3")
13. Gas jar
14. Glass rod
15. Test Tubes (12 x 100, 5 x 5 x 8)
16. Evaporating dish
17. TLC Unit
18. Measuring cylinder
19. Thiele's tubes
20. Fusion Tube
21. Capillary tube
22. Stopper bottle
23. Thermometer (1/10°, 360°)
24. Water condenser
25. Distillation flask (100 ml/ 250 ml)
26. Titration tiles.
27. Asbestos sheet.
28. Desiccators
29. Clay pipe triangle

Iron & Wooden ware:

1. Burners
2. Tripod stand
3. Iron stand
4. wire gauze
5. Burette stand
6. Test tube stand
7. Pair of tongs
8. Test tube holder

9. Spatula

10. Copper foil

Chemicals: All the chemicals required for experiments are mentioned in the syllabus.

Others:

1. Filter papers (Kalpi)

2. Whatman Filter paper No. 1, 40, 41 and 42.

Lab Safety Precautions / Measures in Chemistry Laboratory:

Part-I: Personal Precautions-

1. All personnel must wear safety Goggles at all times.
2. Must wear the Lab. Aprons / Lab jacket and proper shoes.
3. Except in emergency, an over-hurried activity is forbidden.
4. Fume cupboard must be used whenever necessary.
5. Eating, Drinking and Smoking in the laboratories strictly forbidden.

Part-II: Use of safety and Emergency Equipments –

1. First aid kits.
2. Sand Bucket.
3. Fire extinguishers (dry chemical and carbon dioxide extinguisher).
4. Chemical storage cabinet with proper ventilation.
5. Material safety data sheets
6. Management of local exhaust system and fume hoods.
7. Sign in register if using instruments.

B.Sc. III (NEP 1.0)
Revised Syllabus from June 2024 onwards
Semester V & VI Nature of theory Question paper

Total Marks 40

N.B. The question paper should cover the entire syllabus. Marks allotted to questions should be in proportion to the lectures allotted to respective units.

Papers Semester V: IX- DSE-E5, X-DSE-E6, XI- DSE-E7, XII- DSE-E8,
Semester VI: XIII- DSE-F5, XIV-DSE-F6, XV-DSE-F7 and XVI- DSE-F8

Q.1 Choose the correct alternative and rewrite the sentence again

8

- 1)
- 2)
- 3)
- 4)
- 5)
- 6)
- 7)
- 8)

Q.2. Attempt any TWO of the following (Out of Three)

16

- 1)
- 2)
- 3)

Q.3. Answer any FOUR of the following (Out of Six)

16

- 1)
 - 2)
 - 3)
 - 4)
 - 5)
 - 6)
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